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Study Guide

Stoichiometry

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**Answers**

The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

## **VIBRATIONS AND WAVES**

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Questions and Answers

Test your

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understanding with practice problems and step-by-step solutions. Browse through all study tools. How many moles of  $K^+$  ions are present in  $1.67 \times 10^2$  mL...

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Stoichiometry. The atomic ratios in each compound are also the relative number of atomic mass units of

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Answer Key  
its elements. The first example is nitrous oxide ( $N_2O$ ), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and atomic masses.

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### Stoichiometry

RHS - Mr. Moss 1.

Define the following: a.

Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction. b.

Mole -The SI unit used to measure the amount of a substance that contains  $6.02 \times 10^{23}$  atoms of that substance.

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**Stoichiometry Study  
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RHS Mr. Moss**

Stoichiometry is designed to focus on the numbers behind the chemicals involved to ensure that not only are there proper quantities of the reactants, but the product can be measured as well to make...

**Solved: How does  
stoichiometry work?**

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**Calculations:**

**Problems |**

**SparkNotes**

Stoichiometry Example

$C_6H_6 + Br_2 \rightarrow C_6H_5Br + HBr$

Benzene

( $C_6H_6$ ) reacts with

Bromine to produce

bromobenzene ( $C_6H_5Br$ )

and hydrobromic

acid. If 30. g of

benzene reacts with 65

g of bromine and

produces 56.7 g of

bromobenzene, what is

the percent yield of the

reaction? 30.g 65 g

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56.7 g  
Answers

**Chapter 3**  
**Stoichiometry -**  
**Chemistry**

Stoichiometry is the tool for answering these questions.

Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

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### Stoichiometry

Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

#### **Chapter 11: Stoichiometry**

Stoichiometry MIXED  
REVIEW SHORT

ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  4 a. What is



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Answers  
the value of the coefficient  $x$  in this equation? 40.07 g/mol  
b. What is the molar mass of  $C_3H_4$ ? 2 mol  
 $O_2$ :1 mol  $H_2O$  c. What is the mole ratio of  $O_2$  to  $H_2O$  in the above equation? 0.20 mol d.

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**Guided Reading**

**Answer Key**

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Questions. 1. Copper has two stable isotopes  $^{63}_{29}\text{Cu}$  (atomic mass = 62.93 amu; 69.09% natural abundance) and  $^{65}_{29}\text{Cu}$  (atomic mass = 64.9278 amu; 30.91 % natural abundance). Calculate the average atomic mass of copper. 2. How many atoms of sulfur are in 25.1 g of S? 3. Calculate the

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Answers  
molecular mass (in amu) of caffeine (C 8 H 10 N 4 O 2). 4.

## **Stoichiometry**

### **Questions | Shmoop**

Chemistry. From aluminum to xenon, we explain the properties and composition of the substances that make up all matter.

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